B. Tech. 2nd Semester G-Scheme (Common for all branches) Examination, May-2019

CHEMISTRY-I

Paper-BSC-CH-101-G

Time allowed: 3 hours?

[Maximum marks: 75]

Note: Attempt five questions in all, selecting at least one question from each unit. Question No. 1 is compulsory. All questions carry equal marks.

- What is electronegativity?
 - Name the factors that govern the magnitude of crystal field splitting.
 - Distinguish between configuration and conformation.
 - What are the limitations of van der Waal's equation?
 - Define hardness of water. (e)

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What is Bathochromic and Hypsochromic shift? (f)

6×2.5=15

Unit-I

Write schrodinger wave equation for hydrogen. 2. What are the various parameters used in the equation? What are radial and angular wave functions.

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- Draw MO energy level diagram of NO molecule and predict its Bond order and magnetic properties.
- 3. What is Ionization energy? On what factor it depends? How does the ionization energy of the elements vary in a group and period.
 - Explain the polarization and polarizing power. 7

Unit-II

- Draw various conformations of n-butane and compare their stability.
 - Distinguish between enantiomers diastereomers with suitable example.
- 5. What are the main types of organic reaction? Explain substitution reactions giving suitable example. 10
 - Give the synthesis of Aspirin drug.

Unit-III

- How is the hardness of a water sample estimated 6. by EDTA method? Describe the principle involved and the procedure. 10
 - What are critical constants? Derive relationship between them.

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- (a) Describe the zeolite process used for the softening of water.
 - (b) Explain the mechanism of the following:
 - (i) Pitting corrosion
 - (ii) Galvanic corrosion

Unit-IV

- 8. (a) Discuss the applications of IR spectroscopy. 9
 - (b) What is flame photometry? Describe its principle and applications.
- 9. (a) What is chemical shift and on what factors it depends? Explain.
 - (b) Write note on Magnetic Resonance Imaging. 7

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